Gas phase stabiliser-free production of hydrogen peroxide using supported gold–palladium catalysts†


Hydrogen peroxide synthesis from hydrogen and oxygen in the gas phase is postulated to be a key reaction step in the gas phase epoxidation of propene using gold–titanium silicate catalysts. During this process H₂O₂ is consumed in a secondary step to oxidise an organic molecule so is typically not observed as a reaction product. We demonstrate that using AuPd nanoparticles, which are known to have high H₂O₂ synthesis rates in the liquid phase, it is possible to not only oxidise organic molecules in the gas phase but to detect H₂O₂ for the first time as a reaction product in both a fixed bed reactor and a pulsed Temporal Analysis of Products (TAP) reactor without stabilisers present in the gas feed. This observation opens up possibility of synthesising H₂O₂ directly using a gas phase reaction.

Introduction

The direct synthesis of hydrogen peroxide (H₂O₂) from molecular H₂ and O₂ in the liquid phase is an active area of research in the field of heterogeneous catalysts. Catalysts based on Pd and Au–Pd nanoparticles supported on carbon and oxide materials as well as homogeneous gold/palladium systems show interesting properties and have been widely studied in many solvent systems. The direct combination of H₂ and O₂ to form H₂O₂ is carried out with a view to both synthesising viable concentrations of H₂O₂ and utilising synthesised H₂O₂ to carry out in situ oxidation reactions, including the epoxidation of propene.

One of the major problems associated with the vapour phase oxidation of propene to propene oxide (PO) using molecular oxygen is low reaction selectivity. The formation of the epoxide requires the electrophilic addition of an oxygen intermediate to the carbon–carbon double bond, however, propene can also be easily activated via the formation of allylic species which leads to non-selective oxidation. The addition of H₂ to O₂ streams as a sacrificial reductant, to synthesise H₂O₂ or peroxy species in situ, permits the activation of dioxygen at temperatures that are typically much lower than required to activate oxygen alone. Haruta and co-workers were the first to demonstrate that highly dispersed Au/TiO₂ catalysts show an extraordinary selectivity in the oxidation of propene to the corresponding epoxide (>99%), using a combination of H₂ and O₂ as oxidant. H₂ was added as a sacrificial reductant which permits the activation of dioxygen at relatively low temperatures (303–393 K) therefore allowing selective oxidation of propene to propene oxide.

Since TS-1 has been shown to be a selective material for the epoxidation of propene with H₂O₂ as the oxidising species the majority of the early studies used this as a catalyst support. Haruta and co-workers found that Au supported on TS-1 gave high selectivity to propanal above 100 °C however below 100 °C over 90% selectivity to PO could be achieved. In addition Moulijn and others demonstrated that Au/TS-1 catalysts were very stable and could be very selective to the formation of propene oxide. Mechanistic studies showed the important role of the gold nanoparticles in establishing a bidentate propoxy species as an intermediate. It is postulated in many mechanistic studies that H₂O₂ is formed on the Au nanoparticle, rather than a bound peroxy species, which is then capable of either desorbing/adsorbing or migrating on the surface to the tetrahedral Ti site in order to carry out the selective oxidation. In many studies H₂O₂ itself is not observed or analysed for in the reaction mixtures.

Examples in the patent literature report the direct combination of H₂ and O₂ in the gas phase at elevated pressures and temperature, however, these studies include the presence of acid and halides in the gas phase as stabilisers. In this study we aim to utilise catalysts comprising supported AuPd nanoparticles, which have been extensively shown to have a higher H₂O₂ synthesis rate than monometallic Au, to investigate if H₂O₂ can be formed and is able to desorb from the catalyst in the gas phase at atmospheric pressure, without adding stabilisers. Designing catalysts for the synthesis of H₂O₂ as part of...
a gas phase process is an important research target. There are
umerous potential advantages for the synthesis of H₂O₂ in the
gas phase over the current approaches using combined liquid
gas phase reactants. A convenient gas phase process allows
small-scale applications, and this permits simpler process
design if the H₂O₂ is used in a subsequent oxidation reaction. In
contrast the liquid phase suffers from the disadvantage of low
concentrations of O₂ and H₂ which limits the rate of reaction; in
addition there is the potential for leaching of catalyst com-
ponents. Additionally, using a gas phase reaction permits more
facile fundamental research.

Experimental

Supported bimetallic catalysts were prepared by a previously
reported wet impregnation of the appropriate catalyst support
with solutions of PdCl₂ and HAuCl₄. A typical preparation for 1 g
of 2.5% Pd/2.5% Au/TiO₂ was carried out as follows: PdCl₂
(0.0416 g) was added to an aqueous solution of HAuCl₄ (2.04 ml,
12.25 g Au/1000 ml) and heated to 80 °C with stirring and left
until the PdCl₂ had completely dissolved. TiO₂ (0.95 g Degussa
P25) was then added to the solution and the water allowed to
evaporate until the mixture formed a paste. The samples were
dried at 110 °C for 16 h and then calcined in static air at 400 °C
for 3 h with a ramp rate of 20 °C min⁻¹.

Catalytic reactions were carried out in a custom laboratory
scale fixed bed reactor with 1/2 inch diameter PTFE tube. The
reactor bed, typically 50 mg unless stated, was suspended
vertically in an oven by glass wool above and below. A low
reactor bed length/reactor diameter was used to help prevent
back pressure issues. Reaction temperature (40–100 °C) was
stabilised before a flow of 2% H₂ in synthetic air was introduced
(50–200 ml min⁻¹). Reactions were typically carried out for 16 h
using dreschel bottles in an ice bath to condense liquid prod-
ucts from the gas phase. After reaction the presence of H₂O₂ was
determined by titrating aliquots of reaction solution with
acidified Ce(SO₄)₂ solution in the presence of ferroin as an in-
dicator. The Ce(SO₄)₂ solution was standardised against
(NH₄)₂Fe(SO₄)₂·6H₂O using ferroin as an indicator. H₂O₂
decomposition was measured by comparing the difference in
H₂O₂ concentration before and after the reaction using the
same titration method outlined above. Oxidation experiments
were carried out by bubbling gases through a solution of
2-propanol (99.5% as purchased) which once passed over the
catalyst was condensed in an ice bath. NMR analysis of the
reaction mixtures was then carried out to identify the presence
of oxidation products.

The TAP experiments were performed using a TAP-2 system.
A gas mixture containing 2% H₂ in air was pulsed over the
catalyst from a reservoir at a pressure of 0.6 bar. The reactor was
packed with 0.045 g of powdered catalyst held in place with SiC
inert packing. The gas exiting the reactor bed was detected via
quadrupole mass spectrometer. The response for mass 34
(H₂O₂) and mass 2 (H₂) was monitored over a time of 2 s for each
pulse, giving a curve of intensity against time. The experiment
was carried out at 60 °C, with a train of 20 pulses averaged per
experiment.

Results

Initial experiments were carried out with the laboratory scale
reactor to ensure that any H₂O₂ produced could pass through
the reactor system without significant decomposition occur-
ing. A H₂O₂ vapourisation calibration curve was established by
flowing synthetic air through a 50 wt% H₂O₂ solution and
measuring the moles of H₂O₂ carried in the vapour phase at
various flow rates by condensing the gas phase mixture after the
reactor (Fig. 1). The amounts of H₂O₂ in the gas phase suggest
that the gas is not saturated with H₂O₂ due to the high gas flow
rate used and the low vapour pressure of H₂O₂ at ambient
temperature. Following this reactor tubes made of stainless
steel and PTFE were trialled, stainless steel decomposed over
90% of the H₂O₂ vapourised in the gas phase. When a PTFE tube
was used the number of moles of H₂O₂ carried through the tube
in the gas phase matched the calibration curve in both synthetic
air and 2% H₂ in air as shown in Fig. 1. All further experiments
were carried out with a PTFE reaction tube.

Support materials that had been previously used to support
AuPd nanoparticles for H₂O₂ synthesis were screened for H₂O₂
decomposition activity in the gas phase reactor (50 wt% H₂O₂
145 ml min⁻¹ gas flow corresponding to 2.17 mol min⁻¹ from
calibration in Fig. 1). Table 1 shows that all supports tested
showed significant H₂O₂ decomposition, which is in contrast to
the observed decomposition activity in the liquid phase.⁰⁻¹⁹

This can be attributed to the high surface concentrations of
H₂O₂ passing over the catalyst without the protection of solva-
tion that is available in the liquid phase. Despite these high
observed decomposition rates we focused this study on a 2.5%
Au 2.5% Pd/TiO₂ catalyst to investigate whether oxidative
species could be generated from H₂ and O₂ mixtures on AuPd
particles in the gas phase and if H₂O₂ could be synthesised and
detected despite the high decomposition rates of the catalyst
support materials.

The oxidation of 2-propanol to acetone was used to probe the
presence of oxidative species formed by the catalyst. The gas
phase oxidation of 2-propanol using Au catalysts and oxygen

![Fig. 1](image-url)
Temporal analysis of products (TAP) reactor studies were carried out to investigate whether $\text{H}_2\text{O}_2$ is formed and able to desorb from the catalyst surface. Reactions were carried out at 60 °C by pulsing 2% $\text{H}_2/\text{air}$ over the catalyst bed. Fig. 2 shows the accumulated signals observed for mass 34 ($\text{H}_2\text{O}_2$) and 2 ($\text{H}_2$) during the pulses.

A large $\text{H}_2$ signal is detected at the beginning of each pulse along with a smaller signal at mass 34 corresponding to the mass of $\text{H}_2\text{O}_2$. This result shows that it is possible to generate $\text{H}_2\text{O}_2$ in the gas phase and under the vacuum of the TAP reactor it is possible to desorb the $\text{H}_2\text{O}_2$ from the surface of the catalyst and detect it as a reaction product.

Experiments were also conducted in a continuous fixed bed reactor to test under practical conditions for $\text{H}_2\text{O}_2$ synthesis at various temperatures between 40 and 80 °C using 2% $\text{H}_2$ in air at 60 °C. When no catalyst was present with air or 2% $\text{H}_2$ in air no condensate was seen after the reactor in the cold trap indicating that there was no detectable moisture in the gas feeds. In the presence of 50 mg of catalyst with the synthetic air mixture again no condensate was seen in the cold trap. When 2% $\text{H}_2$ in air was used as the reactant gas, liquid was observed in the cold trap indicating that at 60 °C the AuPd nanoparticles were capable of reacting $\text{H}_2$ and $\text{O}_2$ in the gas phase. On titration it was determined that the solution contained 53 ppm of $\text{H}_2\text{O}_2$. This result demonstrates that when using a 5% AuPd/TiO$_2$ catalyst, that is known to have high $\text{H}_2\text{O}_2$ synthesis rates in the liquid phase at 2 °C, it is possible to produce $\text{H}_2\text{O}_2$ in the gas phase reactor system where the residence time is of the order of seconds at ambient pressure. The result also demonstrates that it is possible for $\text{H}_2\text{O}_2$ to desorb from the catalyst into the gas phase after it is formed.

Fig. 3a shows the effect of increasing the reaction temperature from 40–80 °C in the flow reactor. As the temperature is increased the amount of product condensed in the cold trap after 16 h remained constant, full conversion of the $\text{H}_2$ in the

### Table 1

<table>
<thead>
<tr>
<th>Support</th>
<th>$\text{H}_2\text{O}_2$ decomposed/%</th>
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<tbody>
<tr>
<td>Fe$_2$O$_3$</td>
<td>91</td>
</tr>
<tr>
<td>Carbon</td>
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</tr>
<tr>
<td>CeO$_2$</td>
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<tr>
<td>MgO</td>
<td>68</td>
</tr>
<tr>
<td>TiO$_2$</td>
<td>58</td>
</tr>
<tr>
<td>SiO$_2$</td>
<td>53</td>
</tr>
<tr>
<td>Al$_2$O$_3$</td>
<td>47</td>
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</tbody>
</table>

*a* Reaction conditions: catalyst mass 10 mg, 25 °C, 145 ml min$^{-1}$ gas flow.

has been previously studied by Rossi et al.$^{20}$ who showed that oxidation occurred at 393 K with over 80% selectivity to carbonyl derivatives. Other studies have shown that Au/TiO$_2$ is also active for the reaction using oxygen with light-off curves beginning at 350–400 K.$^{21}$ To avoid any reaction with oxygen alone we conducted gas phase oxidation reactions at 60 °C. Gases were bubbled through a solution of 2-propanol (99.5%), lines exiting the catalyst bed were passed through an ice bath so that remaining 2-propanol and liquid products could be condensed for analysis. NMR analysis of the resulting reaction mixtures was then carried out to identify qualitatively the presence of oxidation products.

Initially reactions were carried out with synthetic air and 2% $\text{H}_2$ in synthetic air without the presence of a catalyst, in both cases no oxidation products were detected by NMR (Table 2), similar results were observed when using bare TiO$_2$. Also, when the catalyst was present with 2-propanol in synthetic air no oxidation products were detected. However, when 2% $\text{H}_2$ in air was used as reactant gas NMR analysis showed the presence of acetone in the reaction solutions condensed after passing over the catalyst. This confirms that, at 60 °C, the AuPd particles are able to activate oxygen in the presence of $\text{H}_2$ to oxidise organic molecules such as 2-propanol. This temperature is lower than reported for oxygen alone indicating that $\text{H}_2$ can initiate the oxidation at lower temperature, probably through the formation of a surface bound – hydroperoxy intermediate or through the formation of free $\text{H}_2\text{O}_2$ which is further activated to carry out the oxidation.

### Table 2

<table>
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<tr>
<th>Entry</th>
<th>Catalyst</th>
<th>Gas mixture</th>
<th>Oxidation to acetone</th>
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<tbody>
<tr>
<td>1</td>
<td>None</td>
<td>Industrial grade air</td>
<td>No</td>
</tr>
<tr>
<td>2</td>
<td>None</td>
<td>2% $\text{H}_2$/air</td>
<td>No</td>
</tr>
<tr>
<td>3</td>
<td>TiO$_2$</td>
<td>Industrial grade air</td>
<td>No</td>
</tr>
<tr>
<td>4</td>
<td>TiO$_2$</td>
<td>2% $\text{H}_2$/air</td>
<td>No</td>
</tr>
<tr>
<td>5</td>
<td>5% AuPd/TiO$_2$</td>
<td>Industrial grade air</td>
<td>No</td>
</tr>
<tr>
<td>6</td>
<td>5% AuPd/TiO$_2$</td>
<td>2% $\text{H}_2$/air</td>
<td>Yes</td>
</tr>
</tbody>
</table>

*a* Reaction conditions: catalyst mass 50 mg, 60 °C, 50 ml min$^{-1}$ gas flow – reaction products qualitatively analysed by NMR.

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**Fig. 2** Temporal analysis of products of reaction between of 2% $\text{H}_2$ in air over 2.5% Au 2.5% Pd/TiO$_2$. Reaction conditions – temperature 60 °C, pulses of 2% $\text{H}_2$ in air over 2.5% Au 2.5% Pd/TiO$_2$. Red line – mass spec signal as mass 2 ($\text{H}_2$), black line – mass spec signal at mass 34 ($\text{H}_2\text{O}_2$), blue line – mass spec signal at mass 150 (baseline signal).
The gas stream was observed across the whole temperature range, explaining why we see no increase in liquid reaction product with increasing temperature.

The amount of H$_2$O$_2$ present in the liquid reaction product formed increased from 23 ppm at 40 °C to 55 ppm at 60 °C before decreasing at temperatures above 60 °C. This indicates that there is a balance between the synthesis/desorption and stability of H$_2$O$_2$ through the catalyst bed. H$_2$O$_2$ synthesis rates over the duration of the experiments were approximately $1 \times 10^{-3}$ mol kg$_{\text{cat}}^{-1}$ h$^{-1}$, due to the high H$_2$ conversion and extremely low selectivity leading to mainly H$_2$O production. The observation that it is possible to detect H$_2$O$_2$ at all suggests that the water produced could be the result of a H$_2$O$_2$ decomposition pathway rather than combustion of hydrogen and oxygen without the intermediate production of H$_2$O$_2$. These rates are significantly lower than reported rates in the liquid phase at elevated pressure under optimum conditions (64 mol kg$_{\text{cat}}^{-1}$ h$^{-1}$).

Fig. 3b shows the effect of varying the total gas flow through the catalyst bed over the 16 h reaction. With increasing gas flow the amount of liquid product observed increased linearly, in agreement with the complete conversion of H$_2$ passing over the catalyst. The concentration of H$_2$O$_2$ in the liquid phase increases over the range studied presumably because the faster gas flows can more effectively strip the H$_2$O$_2$ from the catalyst bed.

As no liquid condensate is seen from the gas phase without the presence of the catalyst we can conclude that the water formed is through the combination of H$_2$ and O$_2$ over the catalyst bed. Despite this, the results show that it is possible to produce and collect H$_2$O$_2$ in the gas phase although at low concentrations. Preventing the subsequent decomposition of H$_2$O$_2$ through the catalyst bed remains a challenge never the less this study shows that the concept of producing H$_2$O$_2$ in the gas phase is can be realised.

Conclusions

We demonstrate for the first time that the H$_2$O$_2$ synthesis rates of catalysts that are well studied in the liquid phase with high pressure gas phase reagents are, in fact, high enough to produce H$_2$O$_2$ in the gas phase at atmospheric pressure. Using TAP analysis we have shown that H$_2$O$_2$ can be synthesised and desorbed from the catalyst surface in the gas phase suggesting that gas phase direct synthesis of H$_2$O$_2$ could be feasible. This observation goes some way to explaining the increased reaction rates observed for propene epoxidation when H$_2$ is added to the reactant gas stream. In this study H$_2$ conversions were high and selectivity towards H$_2$O$_2$ was extremely low due to the nature of the packed bed however detectable amounts of H$_2$O$_2$ were synthesised. Through future reactor design and optimization of conditions, coupled with catalyst improvements a simpler process design might be feasible to produce H$_2$O$_2$; the results enable fundamental investigation of the reaction possible through operando surface sensitive spectroscopies.

Acknowledgements

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Notes and references