A Ball-Milling-Enabled Reformatsky Reaction

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An operationally simple one-jar one-step mechanochemical Reformatsky reaction using in situ generated organozinc intermediates under neat grinding conditions has been developed. Notable features of this reaction protocol are that it requires no solvent, no inert gases, and no pre-activation of the bulk zinc source. The developed process is demonstrated to have good substrate scope (39–82% yield) and is effective irrespective of the initial morphology of the zinc source.

Metal-mediated C–C bond formation is an essential tool in modern organic synthesis. Numerous reactions consisting of metal-mediated nucleophilic addition to electrophiles have been developed for the synthesis of complex organic molecules. However, the generally high basicity and/or nucleophilicity of some organometallic reagents restricts their use in late-stage modification, where sensitive functional groups may already exist in the chemical structure. Conversely, organozinc species represent a class of “mild” organometallic compounds that demonstrate excellent functional group compatibility. Nevertheless, the preparation of organozinc species often requires initial access to more reactive organometallics, which are then transmetalated to give the desired organozinc reagent by metathesis with Zn salts. Alternatively, activated Zn can be used for the oxidative addition into carbon–halogen bonds (Scheme 1A). In general the formation and manipulation of organometallic compounds is not particularly clean or green when considering that solvents often have to be distilled and dried prior to use; inert gases are commonly required, and, in the case of organozinc reagents, the form of the bulk metal can play an important role and chemical additives are typically required to generate the activated zinc species. Recently, we identified mechanochemistry and ball-milling as a tool for the straightforward generation of organozinc species without the requirement for carefully prepared solvents or inert gases. Under these conditions the input of mechanical energy is enough to break down the resilient metal oxide surface and, in the presence of the alkyl/aryl halide, generate the corresponding organozinc species (Scheme 1B).

These organozinc species may then be intercepted by opening the grinding jar, adding both catalyst and coupling partner before then running a telescoped Negishi coupling reaction; such a process is applicable to both sp–sp coupling reactions. We have exploited this concept to carry out the one-jar, one-step preparation and use of organozinc species in a robust mechanochemical Reformatsky reaction (Scheme 1C). The Reformatsky reaction offers excellent potential for the formation of C–C bonds through (1) predictable C–C formation, (2) neutral reaction conditions (in comparison to obtaining the same products through aldol condensation), (3) broad functional group tolerance, and (4) the ability to impart a high degree of stereoccontrol.

The classical Reformatsky reaction, in which β-hydroxy esters are formed by the reaction of aldehydes/ketones with α-halo esters in the presence of metallic zinc, was reported in 1887. Since this seminal work, a variety of latent nucleophiles and electrophiles have been studied and applied in this reaction and it has been routinely used in the synthesis of complex natural products. However, to carry out the Reformatsky reaction, organozinc reagents must be prepared at the point of use; the majority of organozinc reagents are not commercially available. This process can be problematic, owing to the formation of a layer of passivating zinc oxide on the surface of zinc powder, which hampers the formation of organozinc species and requires removal through treatment with chemical additives. Such additives include aqueous acid, iodine, and 1,2-di-

Scheme 1. A) Formation and use of organozinc reagents. B) Previous work: one-jar two-step Negishi coupling. C) This work: one-jar one-step mechanochemical Reformatsky reaction.

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bromoethane,[11] or chlorotrimethylsilane.[12] Highly reactive Rieke zinc is an alternative, but its preparation is nontrivial and requires the reduction of ZnCl₂ with alkali metals (Li, Na, K) and naphthalene.[13] Furthermore, the air-sensitivity of this method renders the green credentials of the Rieke approach poor, and the success of the outcome is highly dependent on the physical form of zinc used. Multiple different forms of zinc are commercially available (see the Supporting Information; Figure S1). Therefore, the focus of this work is on the mechanicochemical activation of zinc, which not only renders the process more operationally simple and cost effective, but also delivers significant improvements to some of the green metrics typically associated with this reaction.[14]

Mechanochemistry has been widely used among the crystal engineering and metal–organic framework communities.[15] Recently, ball milling and other mechanochemical techniques have been explored as methods to complement the synthetic toolkit.[16] Running reactions under mechanochemical conditions not only offers a more sustainable way to carry out solvent-minimized/free reactions but can also lead to decreased reaction times, increased selectivity, or different reaction outcomes when compared to results obtained from solution-based reactions.[17] Herein we describe a green method for the Reformatsky reaction by using the ball mill mechanical activation of elemental zinc in air.

Studies commenced by treating model substrates benzaldehyde (1, 1 mmol) and ethyl 2-bromoacetate (2, 1.2 mmol) with 1.6 equivalents of zinc (20–30 mesh zinc granular) at 30 Hz in a 10 mL grinding jar with a single ball of mass 4 g (Table 1). After 2 hours of grinding, ¹H NMR spectroscopy of the crude mixture (mesitylene as internal standard) confirmed that a 70% yield of ethyl 3-hydroxy-3-phenylpropanoate (3) was produced during the milling process (Table 1, entry 1).

Increasing the amount of zinc to 2 equivalents resulted in an 81% yield of the desired hydroxy ester product 3 (Table 1, entry 2). Further increasing the amount of zinc led to no significant increase in yield (Table 1, entries 3 and 4). Rather than increasing the equivalents of zinc, increasing the amount of ethyl 2-bromoacetate (2) from 1.2 to 2.0 equivalents also led to no real difference in the observed yield (Table 1, entries 5 and 6). With the optimized ratio of reagents in hand, a reaction time study assessed four individual reaction times of 0.5, 1, 1.5, and 2 h (Table 1, entries 2 and 7–9), which indicated that the reaction needs 2 h to afford complete conversion. A control experiment, whereby zinc was omitted from the reaction, returned none of the desired product.

We then applied the optimized conditions to a further 11 commercially available forms of zinc (Figure S1). Although, the zinc forms had various particle sizes, which may lead to differences in the ratio between zinc oxide layer and zinc metal, a fixed mass (2 mmol, 0.130 g) of each sample was employed. Pleasingly, we found that in all cases the mechanochemical Reformatsky reaction was successful irrespective of form of zinc that was used (Table 2). Notably, there appears to be a general trend that the forms with a higher surface area/volume ratio performed better for the Reformatsky reaction under neat ball-milling conditions, this is perhaps contrary to prediction as these metal forms should also contain a higher proportion of zinc oxide.

With the optimized conditions established, application to a small range of 11 carbonyl compounds was investigated for reactivity with ethyl 2-bromoacetate and zinc flakes (ca. 325 mesh) under mechanical grinding (Scheme 2). We found that the mechanochemical Reformatsky reaction demonstrates good functional group tolerance. The highly sterically hindered substrate 2,4,6-trimethylbenzaldehyde afforded the...
corresponding hydroxy ester, 5, in good yield (66%). Halogen-containing aromatic aldehydes were also effective substrates, leading to good yields of the isolated halogen-containing products (Scheme 2, 6–8, 77–82 %). Hydrodehalogenation was not observed under these one-pot reaction conditions. Acetophenone was also a competent electrophile under these conditions, affording the isolated tertiary benzylic alcohol product 10 in 69 % yield. Aliphatic aldehydes offered mixed results. Cyclohexylcarboxaldehyde gave the product in 39 % yield, 2-phenylethyl aldehyde afforded 73 % yield of the corresponding product, and cinnamaldehyde resulted in 61 % yield (Scheme 2).

Eight different latent nucleophiles were also examined in this process with benzaldehyde as the model electrophile. Pleasingly, the corresponding organozinc intermediates could be generated under mechanochemical conditions (Scheme 3) and both methyl and tert-butyl α-halo esters could be used to effectively form the β-hydroxy esters 14 and 15 in good to excellent yields. Reaction with ethyl 2-bromopropionate provided 67 % yield of 16 as a mixture of diastereoisomers (syn/anti = 58:42). Reaction with more stereocongested ethyl α-bromoisobutyrate maintained a good yield of 68 % (17, Scheme 3) and α,α-difluoro-β-hydroxyester 18 could also be prepared by this method. The ball-milling-enabled Reformatsky reaction with ethyl 4-bromobut-2-enoate resulted in 52 % yield of the α-substituted product 19 with moderate diastereoselectivity (syn/anti = 62:38). 2-Bromoacetonitrile also participated in the in situ generation of an organozinc reagent and formed the corresponding Reformatsky product 3-hydroxy-3-phenylpropanenitrile (20). Notably, in all cases explored, no reductive aldehyde coupling (pinacol reaction) was observed.

The applicability of imines as electrophiles was also briefly explored under the developed ball-milling conditions. Besides aldehydes and ketones, nonclassical Reformatsky electrophiles such as azomethines, nitriles, lactones, anhydrides, γ-thiolenecats, and amides can also be used for the Reformatsky reaction.[18] Although not fully optimized, N-benzylideneaniline (22) underwent a mechanochemical Reformatsky reaction (Scheme 4A) to afford the β-amino ester 23 in 48 % yield, alongside a small amount of the corresponding β-lactam 24 (7 % yield).

The convenience of the method is well demonstrated by comparing a solution-based reaction with that under ball-mill-
Conflict of interest

The authors declare no conflict of interest.

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